

For each of the following, find the element that best matches the description.

1. This element has 2 valence electrons in the fifth shell. \_\_\_\_\_
2. This element has 7 shells and one valence electron. \_\_\_\_\_
3. This element needs 1 more valence electron to have a full 4<sup>th</sup> shell \_\_\_\_\_
4. This element needs to lose 1 valence electron in its 3<sup>rd</sup> shell to have a full outer 2<sup>nd</sup> shell \_\_\_\_\_
5. This element has a full valence with only 2 electrons \_\_\_\_\_

For the following, write how many valence electrons and shells the element has.

Element	Valence Electrons	# of Shells
Germanium (Ge)		
Polonium (Po)		
Xenon (Xe)		
Indium (In)		
Astatine (At)		

Answer the following questions.

6. What do all the elements in group 5 have in common?  
\_\_\_\_\_
7. What do all the elements in period 6 have in common?  
\_\_\_\_\_
8. Which element will react like Bromine (Br): Krypton (Kr) or Iodine (I)? Why?  
\_\_\_\_\_